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AP Chemistry Chapter 12 Answers Zumdahl12.49For elementary reactions, the rate law can be written using the coefficients in the balanced equation to determine orders.a. Rate = $k[\text{CH}_3\text{NC}]$ b. Rate = $k[\text{O}_3][\text{NO}]$ c. Rate = $k[\text{O}_3]$ d. Rate = $k[\text{O}_3][\text{O}]$ 12.50The observed rate law for this reaction is: Rate = $k[\text{NO}]^2[\text{H}_2]$. For a mechanism to be plausible,the sum of all the steps must give the overall balanced equatio...

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AP Chemistry Chapter 5 Answers Zumdahl5.27a. $4.8 \text{ atm} \times 760 \text{ mmHg} / 1 \text{ atm} = 3.6 \times 10^3 \text{ mmHg}$. $3.6 \times 10^3 \text{ mmHg} \times 1 \text{ torr} / 1 \text{ mmHg} = 3.6 \times 10^3 \text{ torr}$. $4.8 \text{ atm} \times 1.013 \times 10^5 \text{ Pa} / 1 \text{ atm} = 4.9 \times 10^5 \text{ Pa}$. $4.8 \text{ atm} \times 14.7 \text{ psi} / 1 \text{ atm} = 71 \text{ psi}$ 5.296.5 cm x 10 mm 1 cm = 65 mmHg = 65 torr; 65 torr x 1 atm / 760 torr = $8.6 \times 10^{-2} \text{ atm}$ 6.6 x 10-2 atm x 1.013 x 105 Pa / 1 atm = $8.7 \times 10^4 \text{ Pa}$ 5.31If the levels of Hg in each arm of the m...

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AP Chemistry Chapter 9 Answers Zumdahl9.15H₂O has 2(1) + 6 = 8 valance electronsH₂O has a tetrahedral arrangement of the electron pairs about the O atom that requires sp³ hybridization. Two of the four sp³ hybrid orbitals are used to form bonds to the twohydrogen atoms and the other two sp³ hybrid orbitals hold the two lone pairs of oxygen.The two O-H bonds are formed from overlap of the sp³ hybrid...

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AP Chemistry Chapter 10 Answers Zumdahl10.81At 100.°C (373 K), the vapor pressure of H₂O is 1.00 atm = 760. torr.For water, $\Delta H_{\text{vap}} = 40.7 \text{ kJ mol}^{-1}$. $\ln(P_1/P_2) = \Delta H_{\text{vap}}/R(1/T_2 - 1/T_1)$ $\ln(520. \text{ torr} / 760. \text{ torr}) = 40.7 \times 10^3 \text{ J mol}^{-1} / 8.3145 \text{ J K}^{-1} \text{ mol}^{-1} (1/373 \text{ K} - 1/T_2) - 7.75 \times 10^{-5} = (1/373 \text{ K} - 1/T_2) - 7.75 \times 10^{-5} = 2.68 \times 10^{-3} - 1/T_2$ $1/T_2 = 2.76 \times 10^{-3} - 1/373 \text{ K} = 1/362 \text{ K}$ or 89°C10.83 $\ln(P_1/P_2) = \Delta H_{\text{vap}}/R(1/T_2 - 1/T_1)$ $\ln(836 \text{ torr} / 760 \text{ torr}) = 40.7 \times 10^3 \text{ J mol}^{-1} / 8.3145 \text{ J K}^{-1} \text{ mol}^{-1} (1/373 \text{ K} - 1/T_2) - 7.75 \times 10^{-5} = (1/373 \text{ K} - 1/T_2) - 7.75 \times 10^{-5} = 2.68 \times 10^{-3} - 1/T_2$ $1/T_2 = 2.76 \times 10^{-3} - 1/373 \text{ K} = 1/362 \text{ K}$ or 89°C





















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